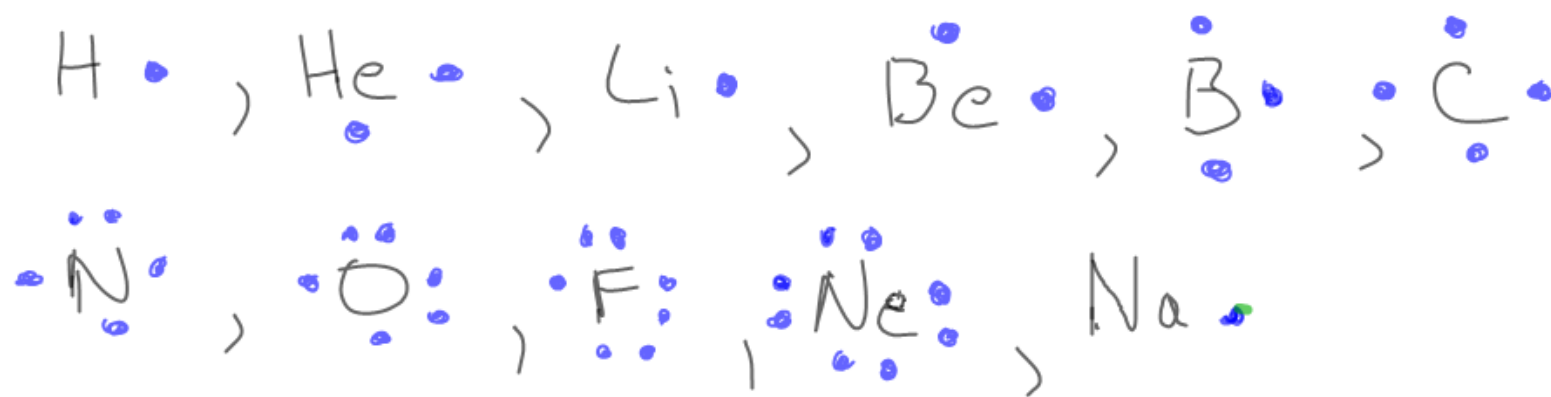


Lewis Electron Dot Diagram



ion - an atom or group of atoms that has an electric charge.

Polyatomic ions - more than one atom that forms a group to make a \oplus or \ominus charge.

ionic bond - attraction between two oppositely charged ions. Usually a metal & non-metal

Covalent bond - sharing electrons - usually non-metals



Silver Nitrate + Sodium Chloride



Silver chloride
precipitate

Naming Compounds

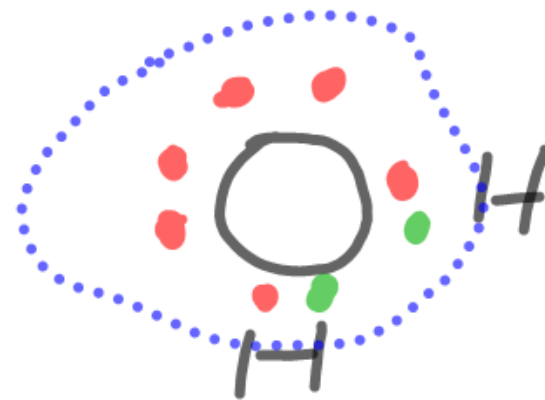
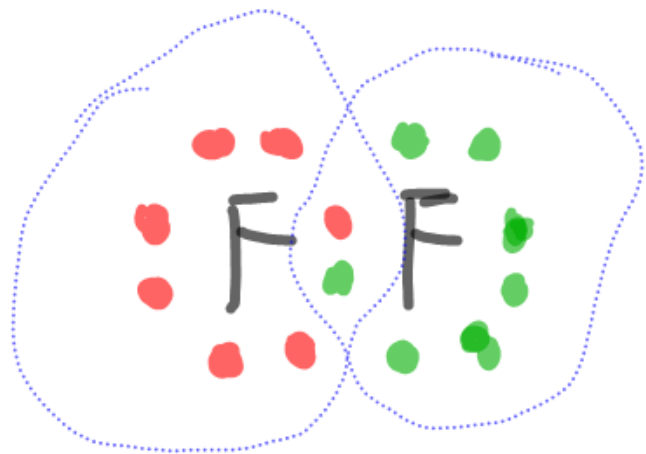
ionic compounds: pos. ion is first, neg. ion 2nd.

if neg. ion is single element - "ide"

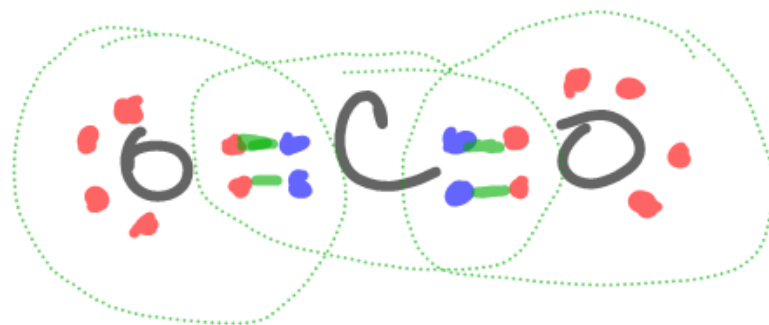
if neg. ion is polyatomic - "ate", "ite"

ionic compounds are generally hard, brittle.
Crystals with ↑ melting pts., conduct
electricity in water.

Covalent bonds - share electrons.
usually between 2
non-metals



Double bonds -



Triple bond -

